

Isolation of a Ru(IV) side-on peroxo intermediate in the water oxidation reaction

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The electrons that nature uses to reduce CO₂ during photosynthesis come from water oxidation at the oxygen-evolving complex of photosystem II. Molecular catalysts have served as models to understand its mechanism, in particular the O–O bond-forming reaction, which is still not fully understood. Here we report a Ru(IV) side-on peroxo complex that serves as a ‘missing link’ for the species that form after the rate-determining O–O bond-forming step. The Ru(IV) side-on peroxo complex (η^2 -1iv–OO) is generated from the isolated Ru(IV) oxo complex (1iv=O) in the presence of an excess of oxidant. The oxidation (IV) and spin state (singlet) of η^2 -1iv–OO were determined by a combination of experimental and theoretical studies. 18O- and 2H-labelling studies evidence the direct evolution of O₂ through the nucleophilic attack of a H₂O molecule on the highly electrophilic metal–oxo species via the formation of η^2 -1iv–OO. These studies demonstrate water nucleophilic attack as a viable mechanism for O–O bond formation, as previously proposed based on indirect evidence.

Water oxidation (WO) to dioxygen is pivotal in natural photosynthesis^{1,2}, and inspires the development of sustainable technologies for the synthesis of renewable fuels and chemicals^{3,4}. Since the first WO complex (the ‘blue dimer’⁵), the focus of molecular catalysts has centred on catalyst design and the characterization of the mechanisms and intermediates that govern the WO reaction^{6–8}. In general, the formal mechanisms considered are the coupling between two radical M–O• (or M–oxyl) species and the acid–base mechanism in which a molecule of water attacks an electrophilic metal oxo species (Fig. 1)^{1,6,9–12}. In both cases, the formation of the O–O bond is usually the step that determines the rate of the reaction, and therefore to facilitate the O–O bond formation should translate into better efficiencies. However, despite indirect proofs for both mechanisms, the evidence for a direct formation of the O–O bond from a M=O moiety is circumstantial^{2,6,11–14}. In part, the challenge is to characterize the intermediates after the O–O bond-forming step, especially when this is the rate-determining step (RDS) of the reaction, which is usually the case. The formation of putative metal–peroxo complexes during catalysis has been a topic of debate over the past decades and remains so^{15–24}.

Since the first report of spectroscopically proposed η^2 -Ru(IV)–O₂ species, [Ru(Mebimpy)(bpy)(η^2 -O₂)]²⁺ (bimpy, bis(imino)pyridine; bpy, 2,2′-bipyridine), based on a Raman band at 1,015 cm^{–1} (ref.¹⁶), several structures have been proposed^{13,18,25–27}. Recently, Garand and co-workers proposed the gas phase generation of 1[Ru(tpy)(bpy)(η^2 -O₂)]²⁺ (tpy, 2,2′;6′,2″-terpyridine) by reacting O₂ with [Ru(tpy)(bpy)]²⁺ (ν (O–O) = 1,150 cm^{–1} and 1,085 cm^{–1} with 18O₂)¹⁸. Yet, compelling evidence for the formation

of M–peroxo species formed under catalytic conditions after the O–O bond formation has not been reported.

Here, we report the isolation and characterization of an elusive Ru(IV) side-on peroxo intermediate η^2 -[Ru(IV)(OO)(L)](PF₆)₂ (η^2 -1iv–OO, L = Py₂ Metacn (tacn, 1,4,7-triazacyclononane) obtained via O–O bond formation from M=O under conditions relevant for catalytic WO with a previously reported [Ru(III)(OH₂)(L)](PF₆)₂ (1) complex²⁷. The η^2 -1iv–OO intermediate was previously postulated to be formed after the O–O bond-formation RDS by computational studies²⁷. Kinetic studies showed a first-order reaction rate versus complex 1 and Ce(IV) and 18O-labelling experiments further proved that the O–O bond formation occurs via water nucleophilic attack (WNA)²⁷.

Results and discussion

Synthesis and characterization of η^2 -1iv–OO. The closed-shell η^2 -1iv–OO was prepared from 1, [Ru(III)(OH)(L)](PF₆)₂ (1iii–OH) or [Ru(IV)(O)(L)](PF₆)₂ (1iv=O) by varying the number of equivalents of added oxidant (NaIO₄, Fig. 2a and Supplementary Section 2). η^2 -1iv–OO was isolated by the reaction of 1iv=O with 3 equiv. NaIO₄ in D₂O at room temperature (r.t.) for three hours, followed by saturation of the solution with ammonium hexafluorophosphate (NH₄PF₆) to give needle-shaped crystals after two days at 5 °C. Two different crystals were analysed to confirm unambiguously the structure, which showed an asymmetric unit that contains one molecule of the metal complex and two counterions (Fig. 2b and Supplementary Section 3). The X-ray crystal structures of η^2 -1iv–OO unequivocally revealed a mononuclear side-on peroxo coordination and, thus, the heptacoordinate configuration of the Ru centre.

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The complex displays a distorted pentagonal bipyramid geometry, which arises from the coordination of the peroxo moiety. The X-ray crystal data refinement clearly shows two oxygen atoms coordinated to the Ru centre with an occupancy of 1.0 for each oxygen atom.

The O–O distances in the isolated crystals were 1.353(10) and 1.367(9) Å and are consistent with the peroxo character of the O–O coordinated moiety, as supported by Raman, X-ray absorption spectroscopy (XAS) and density functional theory (DFT) modeling at the ω B97XD/6-31G(d)/SDD// ω B97XD/cc-pVTZ/SDD (Ru) level of theory (DFT-calculated O–O distance, $d(\text{O–O})_{\text{teor}} = 1.359$ Å for a singlet $\eta^2\text{-1iv-OO}$ complex; Fig. 2)21. The O–O moiety is bound symmetrically to Ru with Ru–O distances and O1–Ru–O2 angle values that fall within those reported for Ru–O2 complexes ($d(\text{O–O}) = 1.36\text{--}1.46$ Å, $d(\text{Ru–O}) = 1.958\text{--}2.040$ Å and $\alpha(\text{O–Ru–O}) = 39.7^\circ$) (ref. 28). The analysis of the diffracted crystal by coldspray high-resolution mass spectrometry (CSI-HRMS) shows a prominent peak at m/z 604.0885 and an isotopic pattern that matches $[[\eta^2\text{-1iv-OO}]2^{++}(\text{PF}_6)^-]_+$ (Supplementary Fig. 35).

DFT calculations were used to estimate the relative stability of $\eta^2\text{-1iv-OO}$ with respect to other isomers. Since previous studies have shown it difficult to accurately describe the O2-binding modes in Ru–O2 (ref. 17), we used two different density functionals to model the possible isomers and compared them with experimental X-ray data. Both functionals, unrestricted ω B97XD and M11, agree that the closed-shell $1[\eta^2\text{-1iv-OO}]$ is the most stable isomer¹⁷.

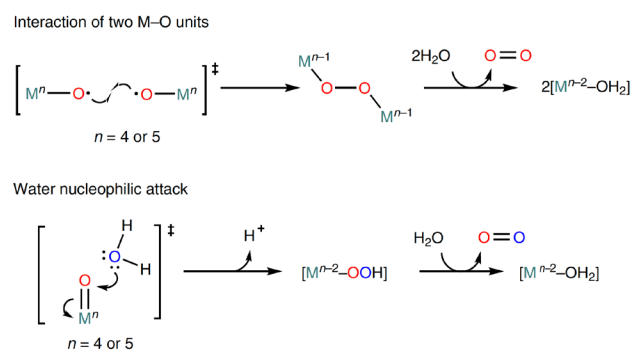


Fig. 1 | O–O bond-forming mechanisms. Formal mechanisms considered in the O–O bond formation.

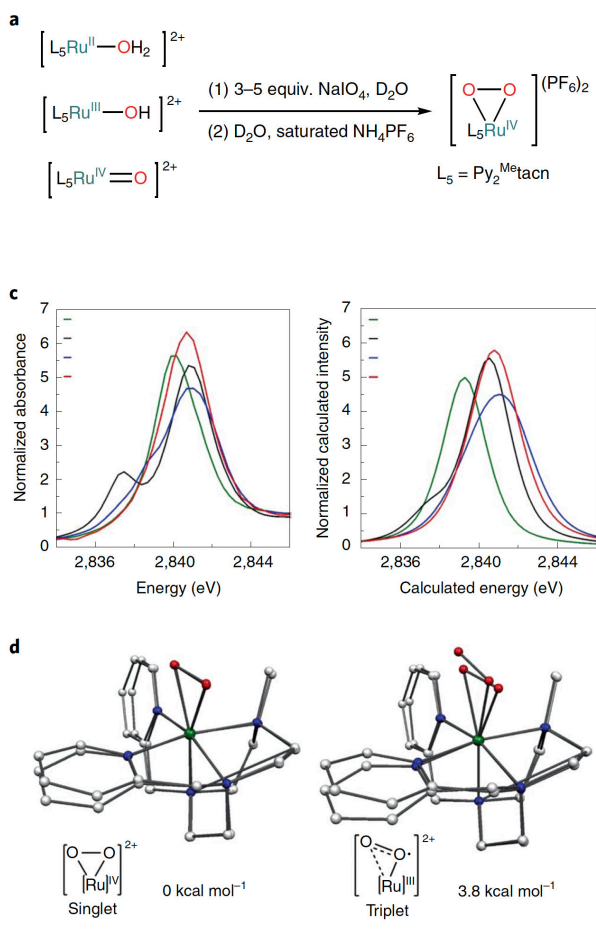
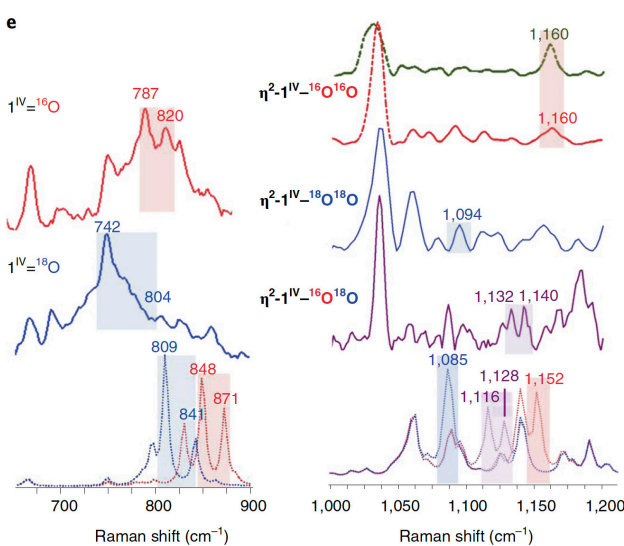


Fig. 2 | Summary of the synthesis and spectroscopic characterization of a closed-shell Ru(IV) side-on peroxo intermediate ($\eta^2\text{-1iv-OO}$). **a**, $\eta^2\text{-1iv-OO}$ was formed from the isolated 1 , 1iii-OH and 1iv=O intermediates, using equivalent synthetic conditions. **b**, ORTEP drawing (thermal ellipsoids are drawn at a 50% probability level) of $[\text{Ruiv}(\text{OO})(\text{Py}_2\text{Metacn})](\text{PF}_6)_{1.5}(\text{IO}_3)_{0.5}$ ($\eta^2\text{-1iv-OO}$). PF_6 anions and hydrogen atoms are omitted for clarity. ORTEP ellipsoids and DFT structures colour code: Ru, green; O, red; C, grey; N, blue. The X-ray and DFT bond distances in the table illustrate a better agreement with the Ru(IV) side-on geometry and singlet electronic structure. **c**, Comparison between the experimental (left) and theoretical (right) Ru L3-edge XAS of 1 (green line), 1iii-OH (black line), 1iv=O (blue line) and $\eta^2\text{-1iv-OO}$ (red line). **d**, The overlap X-ray and

Bond distance (Å)	X-ray	DFT	
	$\eta^2\text{-1iv-OO}$	$1[\eta^2\text{-1iv-OO}]$	$3[\eta^2\text{-1iv-OO}]$
Ru–O ₁	1.996(6)	2.013	2.084
Ru–O ₂	2.004(6)	1.999	2.509
O ₁ –O ₂	1.367(9)	1.359	1.282



DFT-optimized structures of η^2 -1iv-OO in the singlet and the triplet spin states. e, Left: solid-state Raman spectra of the isolated 1iv=O ($\lambda_{exc} = 632.8$ nm, 600 μ W power at r.t.) with 16O (red spectrum) and 18O (blue spectrum) isotopic substitution. Right: solution Raman spectra of the in situ generated η^2 -1iv-OO ($\lambda_{exc} = 457$ nm, 50 mW power at r.t.) from 1iv=O in MilliQ H₂O (green and red spectra generated with cerium(IV) ammonium nitrate (CAN) and NaIO₄, respectively, as the sacrificial oxidants) and from 1iv=18O in H₂ 18O with NaIO₄ (blue spectrum) and D₂O (purple spectrum) (2 mM), in the region of the O–O stretch.

1[η^2 -1iv-OO] is 3.8 kcal mol⁻¹ lower in energy than 3[η^2 -1iii-OO] at the ω B97XD/6-31G(d)/SDD// ω B97XD/cc-pVTZ/SDD (Ru) level of theory. The calculated geometry of 1[η^2 -1iv-OO] overlaps better with the X-ray structure (Fig. 2d and Supplementary Tables 7 and 8). This agrees with the silent electron spin resonance spectrum obtained (Supplementary Fig. 8).

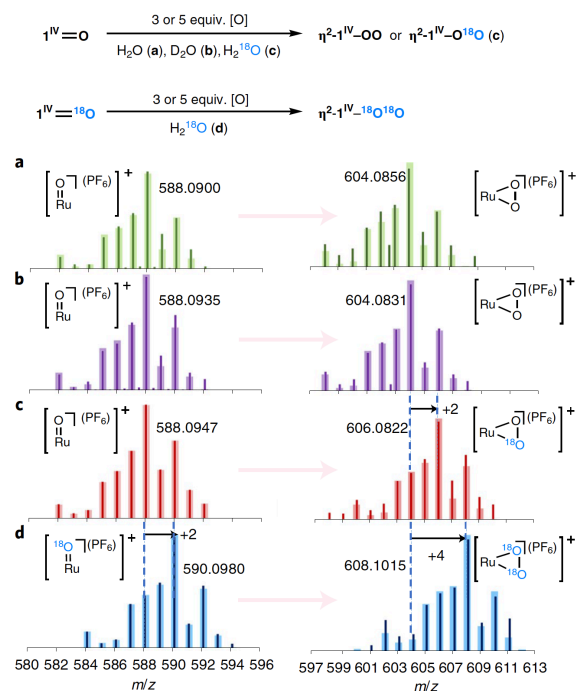


Fig. 3 | CSI-HRMS isotopic labelling experiments evidencing the WNA mechanism. a,b, CSI-HRMS (293 K) of the monocharged [Ru-OO-(PF₆)]⁺ species of η^2 -1iv-OO generated after the addition of NaIO₄ (3 equiv.) to the isolated 1iv=O in MilliQ H₂O (a) and in D₂O (b). c, CSI-HRMS (293 K) of the monocharged [Ru-O18O-(PF₆)]⁺ species of η^2 -1iv-O18O generated after the addition of NaIO₄ (3 equiv.) to the isolated 1iv=O in H₂ 18O. d, CSI-HRMS (293 K) of the monocharged [Ru-O18O-(PF₆)]⁺ species of η^2 -1iv-18O18O generated in H₂ 18O after the addition of NaIO₄ (3 equiv.) to the isolated 1iv=18O. The experimental data (bars) are directly compared with the simulated isotopic pattern (shaded region). Black arrows in the spectra and dashed blue lines indicate the m/z shift on 18O labelling. The observed shift of m/z +2 in H₂ 18O agrees with the incorporation of one 18O into the peroxo moiety in the formed η^2 -1iv-O18O intermediate when starting with 1iv=O after the WNA. The observed shift of m/z +4 in H₂ 18O with respect to that of η^2 -1iv-OO agrees with the two oxygens of the peroxo group being 18O in the final η^2 -1iv-18O18O intermediate when starting with 1iv=18O after the WNA.

The electronic structure of η^2 -1iv-OO was probed by Ru L3-edge XAS (Fig. 2c) and 1, 1iii-OH and 1iv=O were measured as references. η^2 -1iv-OO shows a single intense feature at 2,838.0 eV, consistent with a closely spaced d manifold in a heptacoordinate metal centre. Furthermore, the electronic structure of the η^2 -1iv-OO bond alludes to a peroxo moiety (Supplementary Section 5)²⁹.

The Raman spectrum of η^2 -1iv-OO in water shows a band at 1,160 cm⁻¹ that exhibits an isotope shift of 66 cm⁻¹ on 18O-substitution (1,094 cm⁻¹, Δ 16,18 (DFT-calcd) = 67 cm⁻¹), which was assigned to the O–O stretch (Fig. 2e and Supplementary Section 6). This frequency is within the range of the O–O stretch of related M–O₂ complexes^{16,18–20}. The observed O–O stretch agrees with the formation of a closed-shell heptacoordinate η^2 -1iv-OO intermediate, as supported by the DFT-calculated Raman spectra. The calculated O–O stretch of 1[η^2 -1iv-OO], appears at 1,152 cm⁻¹, whereas an O–O stretch for 3[η^2 -1iii-OO] is observed at 1,297 cm⁻¹ (Supplementary Figs. 26–28)^{16–18,25}. In contrast, the Raman spectrum of 1iv=O shows a Fermi doublet at 787 and 820 cm⁻¹ that exhibits an isotope shift of 45 and 50 cm⁻¹, respectively, on 18O substitution, and so was assigned to the Ru–O stretch (Fig. 2e and Supplementary Figs. 22–25).

Study of the O–O bond formation by isotopic labelling.

The in situ formation of η^2 -1iv-OO from the isolated 1iv=O was investigated by solution Raman and CSI-HRMS labelling studies in MilliQ H₂O, D₂O and H₂ 18O (Supplementary Sections 6 and 7). As the oxo group of 1iv=O does not exchange in H₂ 18O to form 1iv=18O within the experimental time and set-up conditions²⁷, 18O-labelling experiments can be used to track the O–O bond formation event by analysing the nature of the resulting η^2 -1iv-OO. Solution Raman measurements of the in situ generated η^2 -1iv-OO intermediate from the oxidation of 1iv=O by NaIO₄ and Ceiv in MilliQ H₂O, D₂O or H₂ 18O agreed with the formation of a side-on coordinated η^2 -1iv-OO intermediate. A redshift and a split of the O–O stretch band associated with the side-on η^2 -1iv-OO intermediate was observed due to the incorporation of one 18O into the final η^2 -1iv-O18O (two bands at –28 and –20 cm⁻¹ with respect to the single band observed in η^2 -1iv-OO) or η^2 -1iv-18O18O (–66 cm⁻¹) when it was generated from the isolated 1iv=O or 1iv=18O in H₂ 18O, respectively (Fig. 2e and Supplementary Figs. 26–32). Likewise, CSI-HRMS monitoring of a solution that contained 1iv=O (2 mM) and NaIO₄ (5 equiv.) in MilliQ H₂O (or D₂O) showed the formation of a prominent peak at m/z 634.0100 and a less intense one at m/z 604.0885, isotopic patterns associated with [[η^2 -1iv-OO]₂]²⁺(IO₃)⁻ and [[η^2 -1iv-OO]₂]²⁺(PF₆)⁻], respectively (Fig. 3 and Supplementary Figs. 50–53), which rules out any H/D exchange during the process. The intensity of the [[η^2 -1iv-OO]₂]²⁺(PF₆)⁻ peak increases over time due to the accumulation of the intermediate until it reached a plateau (about six hours) due to the formation of [[η^2 -1iv-OO]₂]²⁺ crystals, as observed by microscopy and Raman spectroscopy (Supplementary Fig. 48). This suggests that η^2 -1iv-OO could be isolated under cold-saturated conditions due to its in-

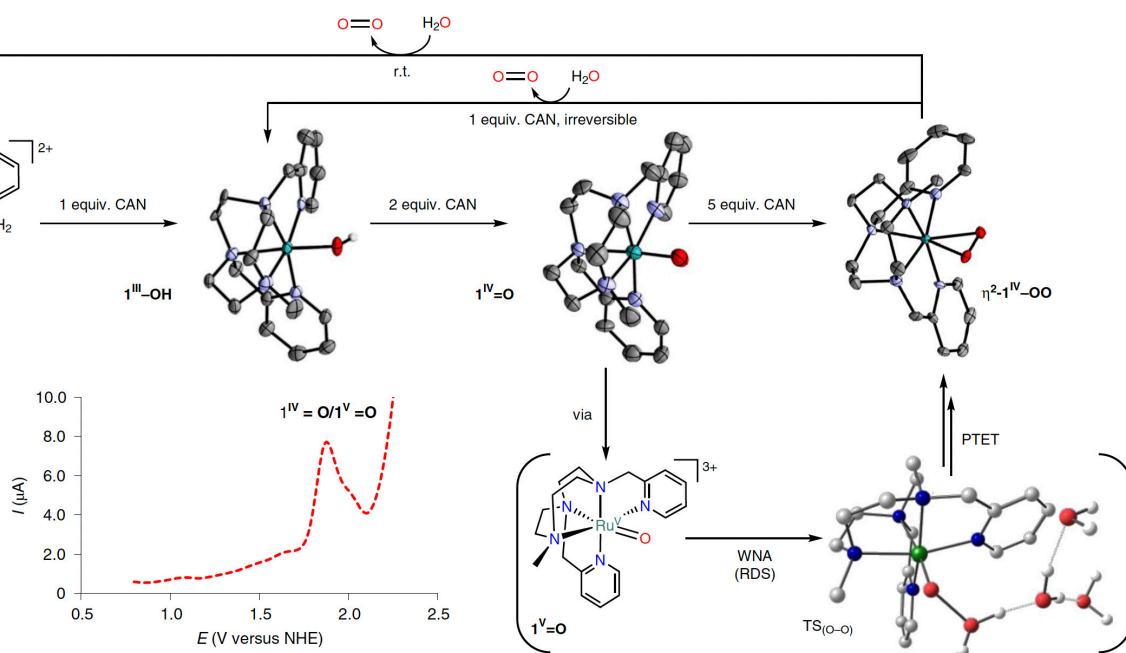


Fig. 4 | Summary of the reactivity of the Ru intermediates in relation to the WO catalytic cycle. Formation and reactivity of ruthenium intermediates involved in WO catalysed by complex 1. Inset: square wave voltammogram (SWV) of $1^{IV}=O$ (1 mM) in a TBAPF6 (0.1 M) propylene carbonate solution under an Ar atmosphere. All the isolated intermediates were catalytically competent with equivalent turnover number and turnover frequency values. The mixed compound labelled $\eta^2-1^{IV}-16O18O$ unequivocally confirmed its generation after WNA onto the highly electrophilic $1^{IV}=O$ species. Likewise, dioxygen $16O=18O$ evolved when starting with the $\eta^2-1^{IV}-16O18O$ intermediate, which confirms the direct oxygen evolution from the isolated Ru(IV) side-on peroxo intermediate. The colour code for the ORTEP ellipsoids and DFT structures: Ru, green; O, red; C, grey; H, white; N, blue. PTET, proton transfer electron transfer. TS, transition state.

solubility (Supplementary Table 10). In addition, a new peak at m/z 229.5601, which corresponds to the dicharged $[\eta^2-1^{IV}-OO]^{2+}$ complex, appeared during the reaction time. These peaks upshifted by m/z 8, 2 and 1 with H_2 18O as the solvent (m/z 642.0313, 606.0944 and 230.5649, respectively), and were assigned to the mix-labelled Ru(IV)-peroxo species $[[\eta^2-1^{IV}-O18O]^{2+}+(18O3)^-]^{+}$, $[[\eta^2-1^{IV}-O18O]^{2++}(PF_6)^-]^{+}$ and $[\eta^2-1^{IV}-O18O]^{2+}$, respectively (Fig. 3 and Supplementary Sections 7.3 and 7.4). The same peaks were obtained when starting with $1^{IV}=18O$ (2 mM), after treatment with $NaIO_4$ (3 equiv.) in H_2 16O. When the reaction was performed from $1^{IV}=18O$ in H_2 18O these peaks upshifted by m/z of 10, 4 and 2, respectively, regarding the unlabelled compounds, associated with $[\eta^2-1^{IV}-18O18O]^{2+}$, $[[\eta^2-1^{IV}-18O18O]^{2++}(PF_6)^-]^{+}$ and $[[\eta^2-1^{IV}-18O18O]^{2++}(18O3)^-]^{+}$, respectively (Supplementary Figs. 44–47). Control experiments showed no O–O exchange of the coordinated peroxo moiety on the addition of H_2O to a solution of $\eta^2-1^{IV}-O18O$ in H_2 18O (Supplementary Fig. 56). Moreover, the fact that we obtained analogous results with the single electron transfer Ce^{IV} ($Ce(OTf)_4$) supports the formation on this in-cycle $\eta^2-1^{IV}-OO$ intermediate via WNA (Supplementary Fig. 57).

Mechanistic studies. At this point, a key aspect is to further understand the connection between the isolated $1^{IV}=O$ and $\eta^2-1^{IV}-OO$, as well as its role in the catalytic cycle. First, electrochemical and ultraviolet–visible spectroelectrochemical studies on the isolated $1^{IV}=O$ intermediate show an irreversible oxidation at 1.85 V versus the normal hydrogen electrode (NHE) (1.77 V versus NHE obtained by DFT calculations; Supplemen-

tary Section 10), which can be assigned to the $1^{IV}/1^{V}$ oxo couple (Fig. 4. and Supplementary Section 8). Previously, $1^{IV}=O$ has been suggested as the active species responsible for the O–O bond formation event on the basis of kinetic and computational studies²⁷.

Computational studies to model the full catalytic cycle at the $\omega B97XD/6-31G(d)/SDD$ (Ru)// $\omega B97XD/cc-pVTZ/SDD$ (Ru) level of theory located the O–O bond formation after the redox pre-equilibrium of the $1^{IV}=O$ to $1^{V}=O$ oxidation (Supplementary Section 10). This is in agreement with our previously observed redox dependence of the WO reaction catalysed by complex 1 on the basis of kinetic studies²⁷. Indeed, all the redox processes are exergonic in the presence of an excess of Ce^{IV} (catalytic conditions), except for the oxidation of $1^{IV}=O$ to $1^{V}=O$. Then, the activation energy obtained from the Eyring equation is consistent with that obtained for the WNA O–O bond formation by DFT (26 ± 2 versus 23 kcal mol⁻¹, respectively; see Supplementary Sections 9.3 and 10 for details). Both theory and experiment suggest that the O–O bond formation is the RDS of the reaction, as judged by the negative ΔS^\ddagger value obtained from the Eyring plot (-29 ± 3 versus -32 cal mol⁻¹), in agreement with an associative mechanism³⁰.

Further evidence on the nature of the RDS is the rapid O_2 evolution (less than five seconds) after the addition of Ce^{IV} (1 equiv.) either to the mixed labelled $\eta^2-1^{IV}-18O16O$ or to $\eta^2-1^{IV}-OO$ in water to yield 1 equiv. $^{34}O_2$ or $^{32}O_2$, respectively, and Ru(III) species. Moreover, the O_2 release from $\eta^2-1^{IV}-OO$ after the addition of 1 equiv. Ce^{IV} was more than 30-fold faster than that

when starting from $1iv=O$ in the presence of a 100-fold excess of Ceiv (Supplementary Fig. 67 and Supplementary Scheme 1) 23. This evidence rules out the O_2 release from $\eta^2-1iv-OO$ as the RDS. Accordingly, DFT calculations show a barrierless O_2 release from the dioxygen-adduct $1iii-O_2$ (Supplementary Figs. 83 and 84, and Supplementary Table 14). The calculated transition state for the O_2 release from $\eta^2-1iv-OO$ (>40 kcal mol $^{-1}$) supports its accumulation in solution at a low concentration of sacrificial oxidant (Supplementary Table 14). The spin conversion to release triplet O_2 could provide an explanation for the slow O_2 release in the absence of additional oxidation.

Catalytic tests for WO from isolated intermediates ($1, 1iii-OH$, $1iv=O$ and $\eta^2-1iv-OO$) showed similar initial O_2 turnover frequencies and numbers (Supplementary Table 11 and Supplementary Fig. 68), consistent with $\eta^2-1iv-OO$ being an intermediate of the catalytic cycle. Moreover, the decay of $\eta^2-1iv-OO$ crystals in acetonitrile gives a Ru(ii) species (Fig. 4 and Supplementary Figs. 70 and 71). Additionally, attempts to obtain $\eta^2-1iv-OO$ under O_2 pressure (up to 20 bar) either in acetone or D2O as solvents or using H_2O_2 in the presence of Et3N did not yield the desired product, which rules out that the O_2 formed, or the potentially formed H_2O_2 , during WO catalysis is responsible for the formation of $\eta^2-1iv-OO$ (Supplementary Fig. 72).

Lastly, a microkinetic model elaborated with the theoretical mechanism and fitted with experimental kinetic data shows that the accumulation of $\eta^2-1iv-OO$ is feasible under a low concentration of sacrificial oxidant (Supplementary Section 11). The kinetic model supports the isolation of $\eta^2-1iv-OO$ as a solid in the presence of ammonium hexafluorophosphate as the precipitating agent (Supplementary Figs. 88 and 89). This mathematical model indicates that an intermediate after the RDS can be trapped in a thermodynamic sink, such as a precipitation equilibrium. Altogether, this is compatible with a catalytic cycle in which $1iv=O$ is the resting state and $\eta^2-1iv-OO$ is transiently formed under catalytic conditions after the RDS. In contrast, by proper modification of the reaction conditions, the elusive $\eta^2-1iv-OO$ can be accumulated and isolated as a solid material. Also, the direct tracking of the O–O bond formation by isotopic labelling from $1v=O$ to $\eta^2-1iv-OO$ supports the nucleophilic attack of water on $1v=O$ to form the O–O bond (RDS) (Fig. 4).

Conclusion

In conclusion, peroxo species have long been postulated as a key intermediate after the O–O bond formation via WNA. Nevertheless, it has only been transiently observed. Here we report a crystallized and isolated $\eta^2-[Ruiv(OO)(Py_2\text{Metacn})](PF_6)_2$ complex generated under catalytic conditions from all the previous isolated intermediates of the catalytic cycle, which is formed after the O–O bond formation event. All the spectroscopic characterization is consistent with a side-on coordination of the peroxo moiety to a closed shell Ru(iv), which is in good agreement with our computa-

tional studies. Moreover, the mixed labelled $16O18O$ Ru(iv)-peroxo complex ($\eta^2-1iv-O18O$) formed on the addition of an excess of Ceiv (single electron oxidant) to $1iv=16O$ (intermediate before the RDS) in H_2 $18O$ points to the formation of the O–O bond via a WNA mechanism. These results prove that the WO mechanism can operate with a single metal site; as one of the proposed mechanisms in an oxygen-evolving complex, in which the M=O species undergoes a nucleophilic attack by water to form the O–O bond in the coordination sphere of the metal site. The isolated $\eta^2-1iv-OO$ intermediate is one of the potential missing links after the RDS, but other missing species need to be revealed to have a more complete view of the post-RDS WO catalytic cycle. This discovery clarifies our understanding of the O–O bond formation event by a direct tracking of the species before and after the RDS.

Online content

Any methods, additional references, Nature Research reporting summaries, source data, extended data, supplementary information, acknowledgements, peer review information; details of author contributions and competing interests; and statements of data and code availability are available at <https://doi.org/10.1038/s41557-021-00702-5>.

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Methods

Py₂ Metacn and complex 1 were synthesized according to previously reported procedures²⁷. Isolated intermediates 1iii–OH and 1iv=O were prepared starting from complex 1 with the addition of an oxidant (NaIO₄, 0.5 and 1.5 equiv., respectively) and further precipitation with an ammonium hexafluorophosphate saturated solution in degassed MilliQ water and then cooling to 5 °C. Likewise, intermediate η^2 -1iv–OO was isolated from all the previous intermediates on the addition of different equivalents of oxidant and further slow precipitation at 5 °C. See the Supplementary Information for further synthetic procedure details, physical methods and detailed procedures for the characterization techniques used.

Data availability

The crystallographic data for η^2 -[Ruiv(OO)(Py₂ Metacn)](PF₆)_{1.5}(IO₃)_{0.5}, η^2 -[Ruiv(OO)(Py₂ Metacn)](PF₆)₂ and 1iv=O have been deposited with the Cambridge Crystallographic Data Centre under accession numbers 1944703, 1944703 and 1944705, respectively. The data supporting the findings of the current study are available within the paper and its Supplementary Information.

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Author contributions

J.L.-F. and C.C. directed and conceived this project. C.C. synthesized all the intermediates, performed most of the experimental work, the DFT studies, the kinetic model and wrote the draft of the manuscript. V.M.-D., C.C., F.F. and B.L.-K. performed the synchrotron measurements and V.M.-D. analysed the data. W.R.B. and C.C. performed the Raman measurements and the analyses. V.M.-D. and C.C. performed the electron spin resonance experiments. S.F. performed the studies to calculate the activation parameters. C.C. and F.F. performed the electrochemical studies. N.C. optimized the experimental parameters of CSI-HRMS and, together with C.C., performed the CSI-HRMS studies and analysed the data. J.B.-B. performed the crystallographic analyses of the isolated intermediates. All the authors discussed the results and contributed to the manuscript.

Competing interests

The authors declare no competing interests.